



Course Outcome Summary

Course Information: (SCAPCH) AP Chemistry

Description:	This AP Chemistry course is designed to be the equivalent of the general chemistry course usually taken during the first year of college. For most students, the course enables them to undertake, as a freshman, second year work in the chemistry sequence at their institution or to register in courses in other fields where general chemistry is a prerequisite. This course is structured around the six big ideas articulated in the AP Chemistry curriculum framework provided by the College Board. A special emphasis will be placed on the seven scientific practices, which capture important aspects of the work that scientists engage in, with learning objectives that combine content with inquiry and reasoning skills. AP chemistry is open to all that have completed a year of chemistry who which to take part in a rigorous and academically challenging course.
Instruction Level:	11 th – 12 th grade
Total Credits:	2
Prerequisites:	Successful completion of Biology/Chemistry/Physics
Textbooks:	Zumdahl Zumdahl Chemistry ninth edition 2014

Course Standards:

- The chemical elements are fundamental building materials of matter, and all matter can be understood in terms of arrangements of atoms. These atoms retain their identity in chemical reactions.
- Chemical and physical properties of materials can be explained by the structure of the arrangement of atoms, ions, or molecules and the forces between them.
- Changes in matter involve the rearrangement and/or reorganization of atoms and/or the transfer of electrons.
- Rates of chemical reactions are determined by details of the molecular collisions.
- The laws of thermodynamics describe the essential role of energy and explain and predict the direction of changes in matter.
- Any bond or intermolecular attraction that can be formed can be broken. These two processes are in a dynamic competition, sensitive to initial conditions and external perturbations.

Unit

1. **Review of Stoichiometry and Nomenclature**
2. **Gas Laws**
3. **Kinetics**
4. **Equilibrium**
5. **Thermodynamics and Electrochemistry**
6. **Atomic Structure and Periodic Table**
7. **Bonding**
8. **Solutions and Intermolecular Forces**

Unit Outlines

1. **Review of Stoichiometry and Nomenclature**

Standards:

- The chemical elements are fundamental building materials of matter, and all matter can be understood in terms of arrangements of atoms. These atoms retain their identity in chemical reactions.
- Changes in matter involve the rearrangement and/or reorganization of atoms and/or the transfer of electrons.

Essential Question:

- What are the fundamental materials of matter, and what are properties of these atoms?
- Why do atoms of each element have unique structures arising from interactions between electrons and nuclei?
- How are atoms conserved in physical and chemical processes?
- How are chemical changes represented by a balanced chemical equation that identifies the ratios with which reactants react and products form?
- How can chemical reactions be classified, considering what the reactants are, what the products are, or how they change from one into the other? Classes of chemical reactions include synthesis, decomposition, acid-base, and oxidation-reduction reactions.
- What are the strong electrostatic forces of attraction holding atoms together in a unit?

Essential Knowledge:

- Molecules are composed of specific combinations of atoms; different molecules are composed of combinations of different elements and of combinations of the same elements in differing amounts and proportions.
- Chemical analysis provides a method for determining the relative number of atoms in a substance, which can be used to identify the substance or determine its purity.
- The mole is the fundamental unit for counting numbers of particles on the macroscopic level and allows quantitative connections to be drawn between laboratory experiments, which occur at the macroscopic level, and chemical processes, which occur at the atomic level.

- An early model of the atom stated that all atoms of an element are identical. Mass spectrometry data demonstrate evidence that contradicts this early model.
- Physical and chemical processes can be depicted symbolically; when this is done, the illustration must conserve all atoms of all types.
- Conservation of atoms makes it possible to compute the masses of substances involved in physical and chemical processes. Chemical processes result in the formation of new substances, and the amount of these depends on the number and the types and masses of elements in the reactants, as well as the efficiency of the transformation.
- A chemical change may be represented by a molecular, ionic, or net ionic equation.
- Quantitative information can be derived from stoichiometric calculations that utilize the mole ratios from the balanced chemical equations. The role of stoichiometry in real-world applications is important to note, so that it does not seem to be simply an exercise done only by chemists.
- Synthesis reactions are those in which atoms and/or molecules combine to form a new compound. Decomposition is the reverse of synthesis, a process whereby molecules are decomposed, often by the use of heat.

2. Gas Laws

Standards:

- The chemical elements are fundamental building materials of matter, and all matter can be understood in terms of arrangements of atoms. These atoms retain their identity in chemical reactions.
- The laws of thermodynamics describe the essential role of energy and explain and predict the direction of changes in matter.

Essential Question:

- What are the fundamental materials of matter, and what are properties of these atoms?
- How can matter be described by its physical properties? The physical properties of a substance generally depend on the spacing between the particles (atoms, molecules, ions) that make up the substances and the forces of attraction among them.
- How can forces of attraction between particles (including the noble gases and also different parts of some large molecules) be used in determining many macroscopic properties of a substance, including how the observable physical state changes with temperature?
- How can chemical changes be represented by a balanced chemical equation that identifies the ratios with which reactants react and products form?
- Why do two systems with different temperatures that are in thermal contact exchange energy?
- How is thermal energy transferred from one system to another?

Essential Knowledge:

- Chemical analysis provides a method for determining the relative number of atoms in a substance, which can be used to identify the substance or determine its purity.
- The mole is the fundamental unit for counting numbers of particles on the macroscopic level and allows quantitative connections to be drawn between laboratory experiments, which occur at the macroscopic level, and chemical processes, which occur at the atomic level.
- The gaseous state can be effectively modeled with a mathematical equation relating various macroscopic properties. A gas has neither a definite volume nor a definite shape; because the effects of attractive forces are minimal, we usually assume that the particles move independently.
- Dipole forces result from the attraction among the positive ends and negative ends of polar molecules. Hydrogen bonding is a strong type of dipole-dipole force that exists when very electronegative atoms (N, O, and F) are involved.
- Intermolecular forces play a key role in determining the properties of substances, including biological structures and interactions.
- Quantitative information can be derived from stoichiometric calculations that utilize the mole ratios from the balanced chemical equations. The role of stoichiometry in real-world applications is important to note, so that it does not seem to be simply an exercise done only by chemists.
- Temperature is a measure of the average kinetic energy of atoms and molecules.

3. Kinetics

Standards:

- Rates of chemical reactions are determined by details of the molecular collisions.

Essential Question:

- What are conditions that affect reaction rates, and how can reaction rates be determined?
- How are elementary reactions mediated by collisions between molecules, and why can only collisions having sufficient energy and proper relative orientation of reactants lead to products?
- How can reactions be expressed via a series of elementary reactions?
- What effect does the presence of a catalyst have on reaction rates?

Essential Knowledge:

- The rate of a reaction is influenced by the concentration or pressure of reactants, the phase of the reactants and products, and environmental factors such as temperature and solvent.
- The rate law shows how the rate depends on reactant concentrations.
- The magnitude and temperature dependence of the rate of reaction is contained quantitatively in the rate constant.
- Elementary reactions can be unimolecular or involve collisions between two or more molecules.

- Not all collisions are successful. To get over the activation energy barrier, the colliding species need sufficient energy. Also, the orientations of the reactant molecules during the collision must allow for the rearrangement of reactant bonds to form product bonds.
- A successful collision can be viewed as following a reaction path with an associated energy profile.
- The mechanism of a multistep reaction consists of a series of elementary reactions that add up to the overall reaction.
- In many reactions, the rate is set by the slowest elementary reaction, or rate-limiting step.
- Reaction intermediates, which are formed during the reaction but not present in the overall reaction, play an important role in multistep reactions.
- Catalysts function by lowering the activation energy of an elementary step in a reaction mechanism, and by providing a new and faster reaction mechanism.
- Important classes in catalysis include acid-base catalysis, surface catalysis, and enzyme catalysis.

4. Equilibrium

Standards:

- Any bond or intermolecular attraction that can be formed can be broken. These two processes are in a dynamic competition, sensitive to initial conditions and external perturbations.

Essential Question:

- What are the characteristics of chemical equilibrium?
- How do systems at equilibrium respond to external perturbations, and how does the response change the composition of the system?
- How does chemical equilibrium play an important role in acid-base chemistry and in solubility?

Essential Knowledge:

- In many classes of reactions, it is important to consider both the forward and reverse reaction.
- The current state of a system undergoing a reversible reaction can be characterized by the extent to which reactants have been converted to products. The relative quantities of reaction components are quantitatively described by the reaction quotient, Q .
- When a system is at equilibrium, all macroscopic variables, such as concentrations, partial pressures, and temperature, do not change over time. Equilibrium results from an equality between the rates of the forward and reverse reactions, at which point $Q = K$.
- The magnitude of the equilibrium constant, K , can be used to determine whether the equilibrium lies toward the reactant side or product side.
- Systems at equilibrium respond to disturbances by partially countering the effect of the disturbance (Le Chatelier's principle).
- A disturbance to a system at equilibrium causes Q to differ from K , thereby taking the system out of the original equilibrium state. The system responds by bringing Q back into agreement with K , thereby establishing a new equilibrium state.

- Chemical equilibrium reasoning can be used to describe the proton-transfer reactions of acid-base chemistry.
- The pH is an important characteristic of aqueous solutions that can be controlled with buffers. Comparing pH to pKa allows one to determine the protonation state of a molecule with a labile proton.
- The solubility of a substance can be understood in terms of chemical equilibrium.

5. Thermodynamics and Electrochemistry

Standards:

- The laws of thermodynamics describe the essential role of energy and explain and predict the direction of changes in matter.

Essential Question:

- How is thermal energy transferred from one system to another?
- What is the Law of Conservation of Energy, and how does it apply to the transfer of energy?
- How are chemical and physical processes driven by a decrease in enthalpy or an increase in entropy, or both?
- What are the characteristics of chemical equilibrium?
- How is the equilibrium constant related to temperature and the difference in Gibbs free energy between reactants and products.

Essential Knowledge:

- Net changes in energy for a chemical reaction can be endothermic or exothermic.
- The process of kinetic energy transfer at the particulate scale is referred to in this course as heat transfer, and the spontaneous direction of the transfer is always from a hot to a cold body.
- Energy is transferred between systems either through heat transfer or through one system doing work on the other system.
- When two systems are in contact with each other and are otherwise isolated, the energy that comes out of one system is equal to the energy that goes into the other system. The combined energy of the two systems remains fixed. Energy transfer can occur through either heat exchange or work.
- Chemical systems undergo three main processes that change their energy: heating/cooling, phase transitions, and chemical reactions.
- Calorimetry is an experimental technique that is used to determine the heat exchanged/transferred in a chemical system.
- The net energy change during a reaction is the sum of the energy required to break the bonds in the reactant molecules and the energy released in forming the bonds of the product molecules. The net change in energy may be positive for endothermic reactions where energy is required, or negative for exothermic reactions where energy is released.
- Some physical or chemical processes involve both a decrease in the internal energy of the components ($\Delta H^\circ < 0$) under consideration and an increase in the entropy of those

components ($\Delta S^\circ > 0$). These processes are necessarily “thermodynamically favored” ($\Delta G^\circ < 0$).

- Entropy is a measure of the dispersal of matter and energy.
- When two systems are in contact with each other and are otherwise isolated, the energy that comes out of one system is equal to the energy that goes into the other system. The combined energy of the two systems remains fixed. Energy transfer can occur through either heat exchange or work.
- If a chemical or physical process is not driven by both entropy and enthalpy changes, then the Gibbs free energy change can be used to determine whether the process is thermodynamically favored.
- External sources of energy can be used to drive change in cases where the Gibbs free energy change is positive.
- A thermodynamically favored process may not occur due to kinetic constraints (kinetic vs. thermodynamic control).
- In many classes of reactions, it is important to consider both the forward and reverse reaction.
- The current state of a system undergoing a reversible reaction can be characterized by the extent to which reactants have been converted to products. The relative quantities of reaction components are quantitatively described by the reaction quotient, Q .
- When a system is at equilibrium, all macroscopic variables, such as concentrations, partial pressures, and temperature, do not change over time. Equilibrium results from an equality between the rates of the forward and reverse reactions, at which point $Q = K$.
- The solubility of a substance can be understood in terms of chemical equilibrium.
- When the difference in Gibbs free energy between reactants and products (ΔG°) is much larger than the thermal energy (RT), the equilibrium constant is either very small (for $\Delta G^\circ > 0$) or very large (for $\Delta G^\circ < 0$). When ΔG° is comparable to the thermal energy (RT), the equilibrium constant is near 1.

6. Atomic Structure and Periodic Table

Standards:

- Chemical and physical properties of materials can be explained by the structure of the arrangement of atoms, ions, or molecules and the forces between them.

Essential Question:

- Why do atoms of each element have unique structures arising from interactions between electrons and nuclei?
- What are the subatomic characteristics that cause elements display periodicity in their properties when the elements are organized according to increasing atomic number?
- How can atomic models be used to explain experimental data on collections of atoms?

Essential Knowledge:

- The atom is composed of negatively charged electrons, which can leave the atom, and a positively charged nucleus that is made of protons and neutrons. The attraction of the

electrons to the nucleus is the basis of the structure of the atom. Coulomb's law is qualitatively useful for understanding the structure of the atom.

- The electronic structure of the atom can be described using an electron configuration that reflects the concept of electrons in quantized energy levels or shells; the energetics of the electrons in the atom can be understood by consideration of Coulomb's law.
- Many properties of atoms exhibit periodic trends that are reflective of the periodicity of electronic structure.
- The currently accepted best model of the atom is based on the quantum mechanical model.
- As is the case with all scientific models, any model of the atom is subject to refinement and change in response to new experimental results. In that sense, an atomic model is not regarded as an exact description of the atom, but rather a theoretical construct that fits a set of experimental data.
- The interaction of electromagnetic waves or light with matter is a powerful means to probe the structure of atoms and molecules, and to measure their concentration.
- Calorimetry is an experimental technique that is used to determine the heat exchanged/transferred in a chemical system.

7. Bonding

Standards:

- The chemical elements are fundamental building materials of matter, and all matter can be understood in terms of arrangements of atoms. These atoms retain their identity in chemical reactions.
- Chemical and physical properties of materials can be explained by the structure of the arrangement of atoms, ions, or molecules and the forces between them.
- The laws of thermodynamics describe the essential role of energy and explain and predict the direction of changes in matter.

Essential Question:

- Why do atoms of each element have unique structures arising from interactions between electrons and nuclei?
- What are the subatomic characteristics that cause elements display periodicity in their properties when the elements are organized according to increasing atomic number?
- What are the strong electrostatic forces of attraction holding atoms together in a unit?
- How is energy transferred in the process of breaking bonds and making bonds?

Essential Knowledge:

- The atom is composed of negatively charged electrons, which can leave the atom, and a positively charged nucleus that is made of protons and neutrons. The attraction of the electrons to the nucleus is the basis of the structure of the atom. Coulomb's law is qualitatively useful for understanding the structure of the atom.
- The electronic structure of the atom can be described using an electron configuration that reflects the concept of electrons in quantized energy levels or shells; the energetics of the electrons in the atom can be understood by consideration of Coulomb's law.

- Many properties of atoms exhibit periodic trends that are reflective of the periodicity of electronic structure.
- The interaction of electromagnetic waves or light with matter is a powerful means to probe the structure of atoms and molecules, and to measure their concentration.
- In covalent bonding, electrons are shared between the nuclei of two atoms to form a molecule or polyatomic ion. Electronegativity differences between the two atoms account for the distribution of the shared electrons and the polarity of the bond.
- Ionic bonding results from the net attraction between oppositely charged ions, closely packed together in a crystal lattice.
- The localized electron bonding model describes and predicts molecular geometry using Lewis diagrams and the VSEPR model.
- Ionic solids have high melting points, are brittle, and conduct electricity only when molten or in solution.
- Potential energy is associated with a particular geometric arrangement of atoms or ions and the electrostatic interactions between them.
- The net energy change during a reaction is the sum of the energy required to break the bonds in the reactant molecules and the energy released in forming the bonds of the product molecules. The net change in energy may be positive for endothermic reactions where energy is required, or negative for exothermic reactions where energy is released.

8. Solutions and Intermolecular Forces

Standards:

- The chemical elements are fundamental building materials of matter, and all matter can be understood in terms of arrangements of atoms. These atoms retain their identity in chemical reactions.
- Chemical and physical properties of materials can be explained by the structure of the arrangement of atoms, ions, or molecules and the forces between them.

Essential Question:

- What are the subatomic characteristics that cause elements display periodicity in their properties when the elements are organized according to increasing atomic number?
- How can forces of attraction between particles (including the noble gases and also different parts of some large molecules) be used in determining many macroscopic properties of a substance, including how the observable physical state changes with temperature?
- How can forces of attraction between particles (including the noble gases and also different parts of some large molecules) be used in determining many macroscopic properties of a substance, including how the observable physical state changes with temperature?
- How are the type of chemical bonding and physical properties related for solid substances?
- What are the electrostatic forces that exist between molecules, as well as between atoms or ions, and why does breaking the resultant intermolecular interactions require energy?
- What are the characteristics of chemical equilibrium?

Essential Knowledge:

- The different properties of solids and liquids can be explained by differences in their structures, both at the particulate level and in their supramolecular structures.
- Solutions are homogenous mixtures in which the physical properties are dependent on the concentration of the solute and the strengths of all interactions among the particles of the solutes and solvent.
- London dispersion forces are attractive forces present between all atoms and molecules. London dispersion forces are often the strongest net intermolecular force between large molecules.
- Dipole forces result from the attraction among the positive ends and negative ends of polar molecules. Hydrogen bonding is a strong type of dipole-dipole force that exists when very electronegative atoms (N, O, and F) are involved.
- Intermolecular forces play a key role in determining the properties of substances, including biological structures and interactions.
- Metallic bonding describes an array of positively charged metal cores surrounded by a sea of mobile valence electrons.
- Ionic solids have high melting points, are brittle, and conduct electricity only when molten or in solution.
- Covalent network solids generally have extremely high melting points, are hard, and are thermal insulators. Some conduct electricity.
- Molecular solids with low molecular weight usually have low melting points and are not expected to conduct electricity as solids, in solution, or when molten.
- Chemical systems undergo three main processes that change their energy: heating/cooling, phase transitions, and chemical reactions.
- Potential energy is associated with the interaction of molecules; as molecules draw near each other, they experience an attractive force.
- At the particulate scale, chemical processes can be distinguished from physical processes because chemical bonds can be distinguished from intermolecular interactions.
- Noncovalent and intermolecular interactions play important roles in many biological and polymer systems.
- In many classes of reactions, it is important to consider both the forward and reverse reaction.